

NOMENCLATURE REVIEW

Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

1. BaSO_3
2. $(\text{NH}_4)_3\text{PO}_4$
3. PBr_5
4. MgSO_4
5. CaO
6. H_3PO_4
7. $\text{Na}_2\text{Cr}_2\text{O}_7$
8. MgO
9. SO_3
10. $\text{Cu}(\text{NO}_3)_2$
11. HI
12. N_2O
13. MnO
14. AgNO_3
15. As_2O_5
16. Fe_2O_3
17. HClO
18. N_2O_3
19. HF
20. $\text{H}_2\text{C}_2\text{O}_4$
21. NaHCO_3
22. SiBr_4
23. CuCl_2
24. HNO_2
25. SnO_2
26. BaCrO_4

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

27. hydrobromic acid
28. chromium(III) carbonate
29. magnesium sulfide
30. iodine trichloride
31. lithium hydride
32. ammonium hydroxide
33. calcium chloride
34. hydroselenic acid
35. iron(II) nitride
36. aluminum hydroxide
37. tin(II) fluoride
38. sulfur tetrachloride
39. mercury(II) iodide
40. diphosphorus pentoxide
41. sulfurous acid
42. lead(II) nitrate
43. dihydrogen monoxide
44. sodium oxalate
45. perchloric acid
46. chlorous acid
47. silicon dioxide
48. carbonic acid
49. sodium chlorate
50. xenon hexafluoride
51. nickel^(II) nitrate
52. potassium perchlorate

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Key

NAME THE FOLLOWING COMPOUNDS:

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|---------------------------------------|-------------------------|--------------------------------------|----------------------|
| 1. BaSO_3 | Barium Sulfite | 14. AgNO_3 | Silver Nitrate |
| 2. $(\text{NH}_4)_3\text{PO}_4$ | Ammonium Phosphate | 15. As_2O_5 | Diarsenic Pentoxide |
| 3. PBr_5 | Phosphorus Pentabromide | 16. Fe_2O_3 | Iron (III) Oxide |
| 4. MgSO_4 | Magnesium Sulfate | 17. HClO | Hypochlorous Acid |
| 5. CaO | Calcium Oxide | 18. N_2O_3 | Dinitrogen Trioxide |
| 6. H_3PO_4 | Phosphoric Acid | 19. HF | Hydrofluoric Acid |
| 7. $\text{Na}_2\text{Cr}_2\text{O}_7$ | Sodium Dichromate | 20. $\text{H}_2\text{C}_2\text{O}_4$ | Oxalic Acid |
| 8. MgO | Magnesium Oxide | 21. NaHCO_3 | Sodium Bicarbonate |
| 9. SO_3 | Sulfur Trioxide | 22. SiBr_4 | Silicon Tetrabromide |
| 10. $\text{Cu}(\text{NO}_3)_2$ | Copper (II) Nitrate | 23. CuCl_2 | Copper (II) Chloride |
| 11. HI | Hydroiodic Acid | 24. HNO_2 | Nitrous Acid |
| 12. N_2O | Dinitrogen Monoxide | 25. SnO_2 | Tin (IV) Oxide |
| 13. MnO | Manganese (II) Oxide | 26. BaCrO_4 | Barium Chromate |

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

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|-----------------------------|--|----------------------------|-----------------------------------|
| 27. hydrobromic acid | HBr | 40. diphosphorus pentoxide | P_2O_5 |
| 28. chromium(III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ | 41. sulfurous acid | H_2SO_3 |
| 29. magnesium sulfide | MgS | 42. lead(II) nitrate | $\text{Pb}(\text{NO}_3)_2$ |
| 30. iodine trichloride | ICl_3 | 43. dihydrogen monoxide | H_2O |
| 31. lithium hydride | LiOH LiH | 44. sodium oxalate | $\text{Na}_2\text{C}_2\text{O}_4$ |
| 32. ammonium hydroxide | NH_4OH | 45. perchloric acid | HClO_4 |
| 33. calcium chloride | CaCl_2 | 46. chlorous acid | HClO_2 |
| 34. hydroselenic acid | H_2Se | 47. silicon dioxide | SiO_2 |
| 35. iron(II) nitride | Fe_3N_2 | 48. carbonic acid | H_2CO_3 |
| 36. aluminum hydroxide | $\text{Al}(\text{OH})_3$ | 49. sodium chlorate | NaClO_3 |
| 37. tin(II) fluoride | SnF_2 | 50. xenon hexafluoride | XeF_6 |
| 38. sulfur tetrachloride | SCl_4 | 51. nickel(I) nitrate | NiNO_3 |
| 39. mercury(II) iodide | HgI_2 | 52. potassium perchlorate | KClO_4 |